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ANSWER Answer the following questions in the space provided. 1.

Given the following equation:  $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient  $x$  in this equation? 40.07 g/mol

b. What is the molar mass of  $C_3H_4$ ? 2 mol

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O 2:1 mol H 2O c. What  
is the mole ratio ...

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the left. Show all your

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provided. 1. \_\_\_\_\_ The

actual yield of a

reaction is 22 g and

the theoretical yield is

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25 g. Calculate the percentage yield.

2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of H

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Composition stoichiometry deals with the mass relationships of elements in

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compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.

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ANSWER Answer the following questions in the space provided. 1.

\_\_\_\_\_ When an electron is added to a neutral atom, energy is (a) always absorbed.

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